*Stoichiometry & Limiting Reactant Quiz*

Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Period: \_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. Given: \_\_AgNO3 + \_\_NiCl2 🡪 \_\_AgCl + \_\_Ni(NO3)2 Answer the following:
   1. Determine the limiting reactant when 2 grams of AgNO3 is reacted with 1 grams of NiCl2 ?
   2. Using the information from (a), what would be the mass of Ni(NO3)2 produced?

* 1. How much of the excess reactant is left over after the reaction is done?

1. Consider the reaction              I2O5(g) + 5 CO(g) ------->  5 CO2(g) + I2(g)
   1. 70.0 grams of iodine(V) oxide, I2O5, reacts with 18.0 grams of carbon monoxide, CO. Determine the limiting reactant.
   2. Using the information from (a), what would be the mass of I2 produced?
   3. How much of the excess reactant is left over after the reaction is done?